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Octahedral Ni-nanocluster (Ni_{85}) for Efficient and Selective Reduction of Nitric Oxide (NO) to Nitrogen (N_2)

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Nitric oxide (NO) reduction pathways are systematically studied on a (111) facet of the octahedral nickel (Ni_{85}) nanocluster in the presence/absence of hydrogen. Thermodynamic (reaction free energies) and kinetic (free energy barriers, and temperature dependent reaction rates) parameters are investigated to find out the most favoured reduction pathway for NO reduction. The catalytic activity of the Ni-nanocluster is investigated in greater detail toward the product selectivity (N_2 vs. N_2O vs. NH_3). The previous theoretical (catalyzed by Pt, Pd, Rh and Ir) and experimental reports (catalyzed by Pt, Ag, Pd) show that direct N-O bond dissociation is very much unlikely due to the high-energy barrier but our study shows that the reaction is thermodynamically and kinetically favourable when catalysed by the octahedral Ni-nanocluster. The catalytic activity of the Ni-nanocluster toward NO reduction reaction is very much efficient and selective toward N_2 formation even in the presence of hydrogen. However, N_2O (one of the major by-products) formation is very much unlikely due to the high activation barrier. Our microkinetic analysis shows that even at high hydrogen partial pressures, the catalyst is very much selective toward N_2 formation over NH_3 .

Reduction of NO has attracted considerable attention in recent years due to the environmental pollutions and industrial applications^{1,2}. NO is produced as a by-product during the combustion of hydrocarbons and is responsible for environmental (such as photochemical smog, acid rain, ozone depletion) and biological problems³⁻⁵. Therefore, the removal of NO is highly sought after and a major challenge for researchers. Extensive progress has been made over Pt⁶⁻¹², Pd¹³⁻¹⁵, Rh¹⁶⁻²⁰ and Ir²¹⁻²³ metal surfaces for the catalytic conversion of NO to N_2 . However, the N-O bond dissociation is one of the most important steps and N-O can be dissociated via two pathways (i) direct N-O bond dissociation or (ii) indirect (hydrogenation followed by N-O bond dissociation). Earlier studies⁶⁻²³ reported that direct N-O bond dissociation is not thermodynamically favourable over Pt⁶⁻¹², Pd¹³⁻¹⁵, Rh¹⁶⁻²⁰ and Ir²¹⁻²³ metal surfaces, but possible from their hydrogenated products (NOH, HNO, HNOH, H_2NO), which increases the possibility of unwanted by-products. Therefore, the relative selectivity of N_2 vs. other products (N_2O and NH_3) varies from catalyst to catalyst. The reduction selectivity toward N_2 formation is more dependent on the direct N-O bond dissociation kinetics. Hence the main objective of this work is to design an efficient catalyst to improve the N-O bond dissociation kinetics to increase the N_2 production selectivity.

NO reduction reactions have been studied mainly over precious metal based catalysts. However, considerable efforts have been made to reduce the cost of metal catalysts and among all the low-cost based metal-catalysts, Ni is the most effective catalyst due to the easy availability of valence shell d-electrons²⁴⁻³⁰. Experimental and theoretical studies have been carried out for NO reduction reactions catalysed by Ni-catalysts³¹⁻³⁷, and it is reported that NO undergoes dissociative adsorption over Ni-catalyst³⁸, resulting in the formation of N_2 as a major product. However, the role of hydrogen (reducing agent) is not considered in any of these studies, and thus a vis-à-vis comparison cannot be made between the direct and indirect N-O bond activation processes. Therefore, it is very important to find out whether Ni-catalyst could trigger the direct N-O bond dissociation or favours the indirect N-O bond dissociation as reported for noble metal based catalysts⁶⁻²³.

Apart from different metal based catalysts, the surface morphology and size of the metal catalyst play important roles in the bond scission process, which in turn control the catalytic activity of a metal catalyst. In this context, metal nanoclusters surrounded by well-defined facets show better catalytic activities compared to their bulk surfaces. Metal nanoclusters enclosed by multi-facets are very noble catalysts and are reported to be very

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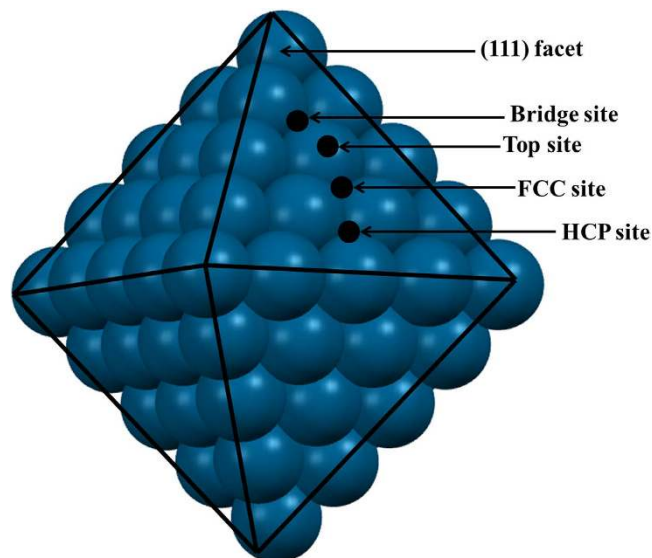


Figure 1. Octahedral Ni₈₅ nanocluster enclosed by eight (111) facets.

promising for electro-catalytic reactions^{39–41} due to the presence of high surface unsaturation. Many theoretical studies have been reported on small size metal nano-clusters^{42–45} to understand their catalytic activities toward nitric oxide reduction. However, the size of the nanoclusters is very important for catalytic reactivity due to their finite-size effects^{46–51} and well-defined facets. These low-coordinated sites have higher d-band energies, which actually increase their reactivity^{42,43}. To the best of our knowledge, there are no reports on NO reduction on a well-defined nanocluster based catalyst. Therefore, to understand the activity of the nanocluster, it is necessary to model a nanocluster surrounded by well-defined facets. The nanocluster with octahedral shape is one of the most stable forms due to its high symmetry^{52,53}. In fact, the Ni(111) surface is mainly observed in the experimentally synthesized Ni-nanoclusters^{54–56}. Hence we have modelled a ~1 nm size of octahedral nickel nanocluster (Ni₈₅) enclosed by well-defined low index facets [eight (111) facets] to understand the NO reduction activity (Fig. 1).

It is also very important to study the complete catalytic process in the presence and absence of hydrogen to understand the underlying NO reduction reaction mechanism to improve the product selectivity. The N₂ production selectivity can be significantly improved if the reduction mechanism goes through the following two steps: direct NO dissociation [$*NO \rightarrow *N + *O$] followed by N₂ formation [$*N + *N \rightarrow *N_2$]. Species are denoted with an asterisk (*) while adsorbed on the surface. But then there are possibilities of by-product formation, such as N₂O formation: $*NO + *N \rightarrow *N_2O$. Koper *et al.*⁵⁷ reported that NH₃ is formed via an indirect N-O bond dissociation process. In that case hydrogenation of NO (formation of $*HNO$ and $*H_2NO$ intermediates) followed by N-O bond dissociation favours the formation of NH₃. Cuesta *et al.*⁵⁸ found $*HNOH$ as the predominant intermediate for NH₃ formation. Previous theoretical studies reported Ir(211) selectively produces N₂, while Pt(211) produces⁸ N₂ and N₂O but they did not consider the role of hydrogen. Carrie *et al.*⁶ studied NO reduction on the Pt (111) surface to understand the N₂O formation as a major product observed experimentally at low temperatures^{59,60}. In spite of several studies, direct N-O bond dissociation is calculated to be the least favourable pathway in the presence/absence of hydrogen gas.

Here we have systematically studied the complete NO reduction pathways on a Ni(111) facet of the nanocluster to find out the most favoured pathway(s) to control the product selectivity. Microkinetic analysis is performed to understand the role of temperature and partial pressure towards the product selectivity. The catalytic activity of the nanocluster is compared with the bulk Ni(111) surface to have a vis-à-vis comparison for the NO reduction reaction. For comparison with a noble metal catalyst, the catalytic activity is compared with that of bulk Pt(111). An attempt has been made to understand the excellent catalytic activity of the nanocluster toward NO reduction and product selectivity compared to any other catalysts reported previously.

Results

Adsorption. Four different catalytic sites (Fig. 1) are present on the (111) facet of the nanocluster: (i) top, (ii) bridge, (iii) hexagonal close packed (hcp) and (iv) face centred cubic (fcc). The relative stabilities of the adsorbed conformers are studied at four possible sites and the most stable conformer is considered for further study.

Intermediates adsorbed on the Ni(111) facet of the nanocluster are presented in Fig. 2. The most preferred binding sites of the intermediates and their respective binding energies are presented (Fig. 2) and compared with previous theoretical studies on Ni(111)^{35,36,61} and Pt(111)⁶ surfaces. The adsorption energies (E_{ad}) for all possible adsorbates are calculated using the following equation:

$$E_{ad} = E_{\text{cluster-species}} - (E_{\text{cluster}} + E_{\text{species}})$$

where $E_{\text{cluster-species}}$ is the total energy of the cluster with adsorbed species, E_{cluster} is the energy of the nickel cluster and E_{species} is the energy of the intermediate species. The binding energies are in very much agreement with

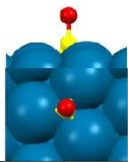
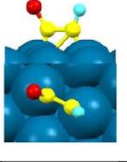
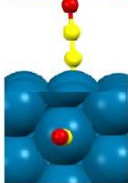
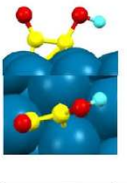
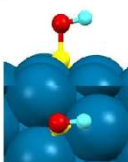
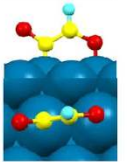
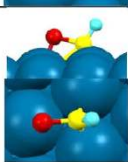
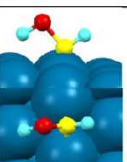
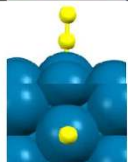
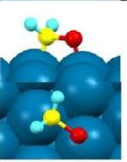
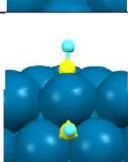
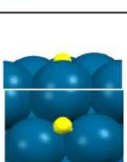
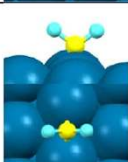
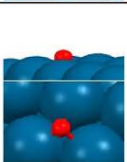
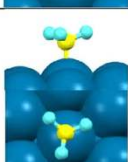
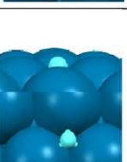
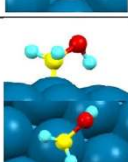
Adsorbate	Optimized structure	Binding energy and site	Adsorbate	Optimized structure	Binding energy and site
*NO		-3.08 (fcc) -2.48 (fcc) ³⁶ -1.88 (fcc) ⁶	*ONNH		-2.93 (top-top) - -1.84 (top-top) ⁶
*N ₂ O		- (top) -0.28 (top) ³⁵ -	*ONNOH		-5.07 (top-br) - -1.84 (top-top) ⁶
*NOH		-4.27 (fcc) - -2.87 (fcc) ⁶	*ONNHO		-2.49 (top-top) - -
*HNO		-4.56 (br-top) - -1.75 (br-top) ⁶	*HNOH		-3.35 (br) - -1.69 (br) ⁶
*N ₂		-0.23 (top) -0.44 (top) ³⁵ -0.12 (top) ⁶	*H ₂ NO		-1.36 (top-top) - -
*NH		-5.05 (fcc) - -4.13 (fcc) ⁶	*N		-6.74 (fcc) -5.49 (fcc) ³⁵ -4.76 (fcc) ⁶
*NH ₂		-3.02 (br) - -2.44 (br) ⁶	*O		-5.02 (fcc) -5.63 (fcc) ³⁵ -4.31 (fcc) ⁶
*NH ₃		-0.51 (top) -0.83 (top) ⁶¹ -0.66 (top) ⁶	*H		-2.85 (fcc) - -2.76 (fcc) ⁶
*H ₂ NOH		- (top-br) - -0.94 (top) ⁶			

Figure 2. Site preference, binding energy (eV) of the most stable adsorbates on the (111) facet of the Ni₈₅ naocluster. The sites preferences and binding energies of the respective adsorbates are compared with previous reports on Pt(111)⁶ and Ni(111)^{35,36,61} surfaces. Yellow, cyan, red, and blue spheres represent N, H, O, and Ni, respectively.

previous studies^{6,7,35–37} though some of our calculated binding energies are higher than the previous reported values. The differences in adsorption energies compared to previous theoretical values could be due to the excellent catalytic activity of the Ni-nanocluster. The adsorption patterns of some of the important intermediates (*NO, *N₂O, *N₂, *NOH, *HNO, *H₂NOH and *NH₃) are discussed and compared to previous experimental and theoretical reports.

***NO.** *NO prefers to be adsorbed at the fcc site (Fig. 2) of Ni(111) through its N atom with an adsorption energy of -3.08 eV. Wu *et al.*³⁵ reported that the *NO molecule prefers to occupy the fcc site of the Ni(111) surface with an adsorption energy of -2.46 eV. It has been experimentally³⁸ as well as theoretically^{35,62} reported that NO prefers to be adsorbed through its N atom as it has more vacant orbitals than an O-atom. The vibrational frequency of the *N-O bond ($1460\text{--}1543\text{ cm}^{-1}$) while adsorbed at the fcc site of Ni(111) surface is characterized³⁸ using RAIRS, EELS and PED, and agrees well with our calculated value of 1530 cm^{-1} . Therefore, our results are in very much agreement with previous reports on binding energy and vibrational frequency^{35–37}.

***N₂O.** *N₂O prefers (Fig. 2) to be adsorbed at a top site of the Ni(111) facet. A theoretical study⁶³ on the (111) terrace of the high index Ni(755) surface reported a \angle N-N-O bond angle of 180° and Ni-N, N-N and N-O bond distances of 1.84, 1.15 and 1.20 Å, respectively. They are in very much agreement with our calculated (Fig. 2) \angle N-N-O bond angle of 179.66° and Ni-N, N-N and N-O bond distances of 1.83, 1.14, 1.20 Å respectively. Using XPS and TPD studies, they reported that *N₂O weakly adsorbed on the Ni(111) surface through its N-atom^{64,65}. Thus, the calculated structural parameters are in very much agreement with the previous report³⁵. This suggests that our study on the Ni(111) facet is in very close agreement with previous experimental reports on the high index Ni(755) surface^{64,65}.

***NOH and *HNO.** In the presence of hydrogen, *NO can react with *H to form *NOH and *HNO. *NOH adsorbs strongly at the fcc-site (Fig. 2) through its N atom with adsorption energy of -4.27 eV. On the other hand, *HNO binds in a di-sigma manner (Fig. 2) with an adsorption energy of -4.56 eV. Their binding preferences are very much in agreement with previous findings^{6,7,13}.

***N₂.** *N₂ binds perpendicular to the top site of the (111) facet of the nanocluster (Fig. 2) with adsorption energy of -0.23 eV. The calculated Ni-N and N-N bond distances are 1.78 Å and 1.12 Å respectively. The N-N bond distance is slightly lengthened while adsorbed on the surface compared to the N-N distance (1.10 Å) in the N₂ gas molecule, suggesting that the interaction is very weak between the surface and gas molecule. Previous studies^{35,36} on the Ni(111) surface reported that the top site is the most favoured position for N₂ adsorption. Using infrared reflection absorption spectroscopy, Yoshinobu *et al.*⁶⁶ reported that *N₂ adsorbs perpendicular on the Ni(111) surface with an adsorption energy of -0.36 eV.

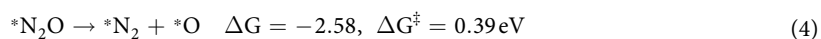
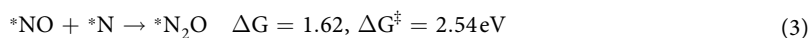
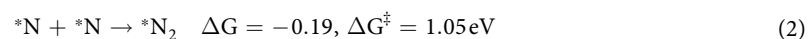
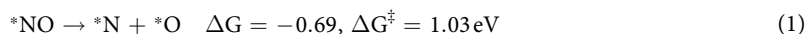
***NH_x (x = 1–3).** *NH_x containing species (Fig. 2) can be formed during the course of reaction. These intermediates (*NH, *NH₂ and *NH₃) are most stable at the fcc, bridge and top sites with Ni-N bond distances of 1.81, 1.91 and 2.01 Å respectively (Fig. 2). *NH and *NH₂ bind strongly, whereas *NH₃ binds weakly on the Ni-nanocluster surface. The calculated adsorption energies are -5.05 eV, -3.02 eV and -0.51 eV for *NH, *NH₂ and *NH₃ respectively. Sergey *et al.*⁶¹ reported *NH₃ prefers to bind on the top site of the Ni(111) surface, which agrees well with our finding. The weak adsorption of *NH₃ suggests the easy desorption of NH₃ from the catalyst surface, which in turn reduces the possibility of surface poisoning.

***H₂NOH.** *H₂NOH can be formed via hydrogenation at the N- and O- sites of *HNOH and *H₂NO. *H₂NOH adsorbs on the top-bridge position by its N- (Ni-N = 1.97 Å) and O atoms (Fig. 2).

It is interesting to find out that adsorption energy decreases (Fig. 2) from NO to higher hydrogenated products implying that the interactions are weaker for higher hydrogenated products.

Reaction Free Energy and Activation Barriers. The reaction free energies are calculated from the total energy difference between the products and reactants. Thus negative free energy suggests the exergonic nature of the reaction, whereas positive reaction energy suggests the endergonic nature of the reaction. Activation barriers are calculated from the energy differences between the transition and initial states.

***N₂ and *N₂O Formation.** The direct N-O bond dissociation (Fig. 3) can lead to the formation of important products such as *N₂, and *N₂O.



The direct N-O bond dissociation (step 1) is a downhill process (-0.69 eV) with an activation barrier of 1.03 eV. The *N-O bond distance is 1.22 Å, which is higher than the N-O bond distance (1.15 Å) in the gas (NO) molecule. This suggests that the gas molecule gets activated during the adsorption process. The previously calculated N-O bond dissociation barriers are 1.49 eV³⁶ and 1.58 eV³⁵ on Ni(111) surfaces, which are far higher than

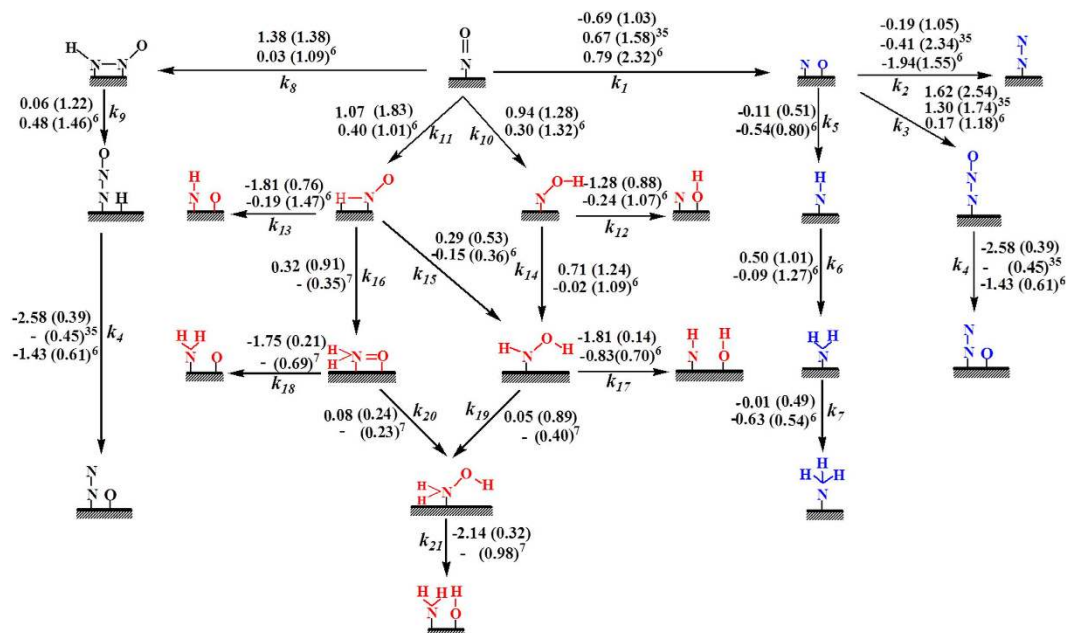


Figure 3. Reaction Scheme. Reaction free energies (eV) and activation barriers (eV, in parenthesis) are presented for all the possible elementary steps for NO reduction over the (111) facet of the Ni_{85} nanocluster. Our calculated respective values are compared with previous reports on NO reduction over $\text{Ni}(111)$ ³⁵ and $\text{Pt}(111)$ ^{6,7} bulk surfaces.

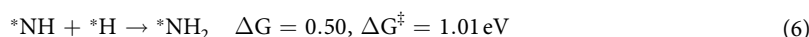
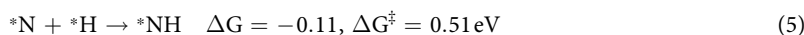
our calculated barrier of 1.03 eV. In fact, the higher *N-O bond distance (1.22 Å), compared to previous report (1.18 Å)³⁷, suggests the excellent catalytic activity of the nanocluster. Earlier reports of N-O bond dissociation barriers on bulk $\text{Pt}(111)$ surfaces are 2.60 eV⁸ and 2.32 eV⁶, respectively, which are considerably higher than our calculated barrier of 1.03 eV. Therefore, Ni_{85} nanocluster is certainly an efficient catalyst for the direct N-O bond activation reaction.

The formation of *N_2 (step 2) from the atomic nitrogens (*N) is exergonic by -0.19 eV with an activation barrier of 1.05 eV. Wu *et al.*³⁵ reported an activation barrier of 2.34 eV for the same step on the bulk $\text{Ni}(111)$ surface. Carrie *et al.*⁶ studied the *N_2 formation on $\text{Pt}(111)$ surface and reported a free energy barrier of 1.55 eV. Therefore, the Ni_{85} nanocluster shows excellent catalytic activity toward *N_2 formation compared to noble metal based catalysts.

*N_2 formation is possible via another pathway where $\text{*N}_2\text{O}$ (step 3) formation takes place, followed by $\text{*N}_2\text{O}$ dissociation (step 4). The $\text{*N}_2\text{O}$ formation is highly endergonic (1.62 eV) with a free energy barrier of 2.54 eV, whereas $\text{*N}_2\text{O}$ dissociation is highly exergonic (-2.58 eV) with an activation barrier of 0.39 eV. Wu *et al.*³⁵ reported the $\text{*N}_2\text{O}$ formation barrier of 1.74 eV (step 3) on the $\text{Ni}(111)$ surface. Carrie *et al.* reported the $\text{*N}_2\text{O}$ formation and dissociation barriers⁶ of 1.18 eV and 0.61 eV on $\text{Pt}(111)$. The calculated barrier for $\text{*N}_2\text{O}$ formation is significantly higher than the previous reports on $\text{Ni}(111)$ and $\text{Pt}(111)$ surfaces.

Wu *et al.*³⁵ studied the first three steps of NO reduction on the $\text{Ni}(111)$ surface as follows: (i) *NO dissociation, (ii) *N_2 and (iii) $\text{*N}_2\text{O}$ formation. The calculated activation barriers are 1.58 eV (*NO dissociation), 2.34 eV (*N_2 formation) and 1.74 eV ($\text{*N}_2\text{O}$ formation) respectively. Similarly Carrie *et al.*⁶, reported activation barriers of 2.32 eV (*NO dissociation), 1.55 eV (*N_2 formation), 1.18 eV ($\text{*N}_2\text{O}$ formation), and 0.61 eV ($\text{*N}_2\text{O}$ dissociation) over $\text{Pt}(111)$. Therefore, the previous reports show $\text{*N}_2\text{O}$ formation is favourable over *N_2 formation, whereas our study shows *N_2 formation is favourable over $\text{*N}_2\text{O}$ formation. However, $\text{*N}_2\text{O}$ dissociation is very favourable (step 4) in all the cases. But on the Ni_{85} cluster's surface, the $\text{*N}_2\text{O}$ formation is very unlikely to happen due to the high activation barrier of 2.54 eV. Therefore, our nanocluster shows excellent catalytic activity toward *N_2 formation, compared to earlier reports^{6,8,35–37}. Hence, the dissociation of the N-O bond followed by N-N bond formation ($\text{*NO} \rightarrow \text{*N} \rightarrow \text{*N}_2$) is the most favourable pathway in the absence of hydrogen.

***NH₃ Formation.**



Hydrogen is used as a reducing agent for NO reduction. The dissociation of hydrogen toward atomic hydrogens ($*H_2 \rightarrow *H + *H$) is highly exergonic (-1.07 eV) with an activation barrier of 0.02 eV. The dissociative adsorption of hydrogen on the Pt (111) bulk surface is reported to be exergonic by -0.87 eV with an activation barrier of 0.00 eV⁶. The activation barrier for hydrogen dissociation on Ni(111) surface is very much comparable with that on the Pt(111) surface. However, the reaction is highly exergonic on the Ni-surface compared to the Pt(111) surface^{25,67}. Therefore, the Ni-nanocluster is a promising candidate for the hydrogenation reaction.

The first hydrogenation on *N (step 5, *NH formation) has an activation barrier of 0.51 eV though the reaction is exergonic by -0.11 eV. In contrast, the second hydrogenation (step 6) for the formation of *NH₂ is endergonic by 0.50 eV with an activation barrier of 1.01 eV. The formation between *NH₂ (step 6) and *N₂ (step 2) can be very competitive as both the reaction steps (Fig. 3) have comparable free energy barriers (1.01 – 1.05 eV). However, the third hydrogenation (step 7, *NH₃ formation) has an activation barrier of 0.49 eV with reaction free energy of -0.01 eV, suggesting that the *NH₃ formation is very much favourable.

Hence, in the presence of hydrogen, *NO dissociation could lead to competition between *N₂ and *NH formation. The calculated activation barriers show that *NH formation (barrier = 0.51 eV) is more favourable than *N₂ formation (barrier = 1.05 eV). However, *NH₂ formation (barrier = 1.01 eV) is not much favourable compared to *N₂ formation (barrier = 1.05 eV). But, our reaction thermodynamics shows, *N₂ formation is highly exergonic (-0.19 eV) compared to *NH₂ formation (0.50 eV). Therefore, *N₂ formation is a spontaneous process where as *NH₂ formation is not thermodynamically favourable. Even the *NH (*NH \rightarrow *N + *H) dissociative barrier (Fig. 3) is calculated to be 0.62 eV with a reaction energy of 0.11 eV, indicating the tendency of *NH to decompose to *N instead of hydrogenation to *NH₂. Therefore *N₂ formation is very much favourable, compared to ammonia formation even in the presence of hydrogen.

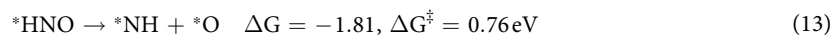
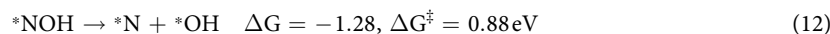
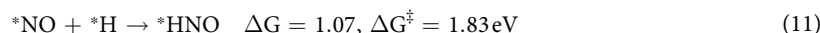
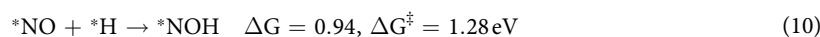
The hydrogenation at *NO on the Pt(111) surface leads to the indirect dissociation of the N-O bond, which in turn reduces the selectivity of *N₂ formation over *NH₃ formation^{68,69}. Surprisingly, we find that the direct N-O bond dissociation barrier is lower compared to the indirect (*N-OH and *HN-O) bond dissociation barrier on the Ni(111) facet of the nanocluster. Therefore, on the Ni-nanocluster surface, *NO can be reduced to *N₂ even in the absence of hydrogen, which minimizes the possibility of unwanted by-products (*NH₃ and *N₂O) formation.

*N₂O can also be formed (Fig. 3) by the direct combination of *NO and *NH (step 8) followed by the dissociation of *ONNH (step 9).



The formation of *ONNH from *NO and *NH (step 8), is highly endergonic (1.38 eV) with an activation barrier of 1.38 eV. However, the dissociation of *ONNH (step 9) leads to the formation of *N₂O, which is mildly endergonic (0.06 eV), with an activation barrier of 1.22 eV. But *N₂O dissociation is a spontaneous process with an activation barrier of 0.39 eV (step 4). Thus, *N₂O formation is very unlikely due to the high activation barrier.

*NO Hydrogenation. First Hydrogenation and Dissociation steps.



In the hydrogenation pathways (Fig. 3), *NO can be hydrogenated to form *NOH and *HNO respectively. Here, we have studied all the hydrogenated intermediates (Fig. 3) to find out the thermodynamic and kinetic feasibility of such reactions.

The *NO hydrogenation reactions (step 10 and 11) are endergonic by 0.94 eV (step 10) and 1.07 eV (step 11) with activation barriers of 1.28 eV (step 10) and 1.83 eV (step 11), respectively. Therefore, hydrogenation at the O atom is more favourable over the N atom. The relative stabilities of the isomers (*NOH and *HNO) show that *NOH is more stable (by 0.15 eV) over *HNO. Previous reports also show that NO undergoes hydrogenation at O⁻⁷ and N- atoms⁶ though favourable at the O atom.

The reaction energies for N-O bond dissociation of the hydrogenated species are -1.28 eV (step 12) and -1.81 eV (step 13) and activation barriers are 0.88 eV and 0.76 eV for *NOH and *HNO respectively. Therefore, hydrogenation on NO (Fig. 3) lowers the N-O bond dissociation barrier. Our study shows that *HNO has a lower N-O bond dissociation barrier than *NOH. As *NH and *N are the two competitive intermediates for *NH₃ and *N₂ formation, therefore N-O dissociation barriers of *HNO and *NOH give an idea about the selectivity of product formation. Carrie *et al.*⁶ reported that the dissociation barrier of the N-O bond is less in *NOH than in *NO on the Pt (111) surface, suggesting that *N₂O is the predominant product. In fact, all the previous studies⁶⁷ show indirect N-O bond dissociation is favourable over direct N-O bond dissociation, whereas we find direct N-O bond dissociation is favourable over indirect N-O bond dissociation. Therefore, the Ni₈₅ nanocluster is an efficient and a selective catalyst for *N₂ formation over *N₂O and *NH₃ formation.

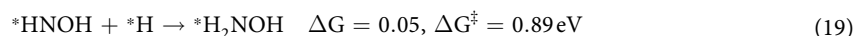
Second Hydrogenation and Dissociation steps.

The $*NOH$ can further undergo hydrogenation for the formation of $*HNOH$ (step 14). It has an activation barrier of 1.24 eV with reaction free energy of 0.71 eV. Similarly $*HNO$ can undergo hydrogenation for the formation of $*HNOH$ (step 15) and $*H_2NO$ (step 16). $*HNOH$ formation is endergonic by 0.29 eV with an activation barrier of 0.53 eV and $*H_2NO$ formation is endergonic by 0.32 eV with an activation barrier of 0.91 eV. Hence, hydrogenation on $*HNO$ is thermodynamically and kinetically favourable over hydrogenation on $*NOH$.

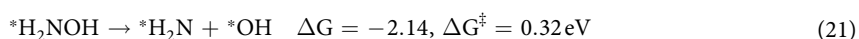


Now, both the hydrogenated intermediates ($*HNOH$ and $*H_2NO$) can undergo N-O bond dissociation to produce $*NH$ and $*NH_2$ respectively (step 17–18). The dissociation reactions (step 17–18) are highly exergonic and have low activation barriers and are easy to dissociate.

Our results show that $*NO$ prefers to be hydrogenated at the O-centre to form $*NOH$ followed by hydrogenation at the N-centre to form $*HNOH$, which support the preference of hydrogenation at the O-centre over the N-centre. The preference of hydrogenation at the O-centre is also found for the first hydrogenation of $*NO$ ($*NO + *H \rightarrow *NOH$). Andre *et al.*⁷ also reported that $*NOH$ and $*HNOH$ are the favourable intermediates over $*HNO$ and $*H_2NO$ intermediates. Therefore, the most favourable pathway for $*NH_3$ formation is $*NO \rightarrow *NOH \rightarrow *HNOH \rightarrow *NH$, whereas the least favourable pathway is $*NO \rightarrow *HNO \rightarrow *HNOH \rightarrow *NH$.

Third Hydrogenation and Dissociation steps.

The third hydrogenation (step 19–20) on $*NO$ can take place for the formation of $*H_2NOH$ and $*H_2NOH$. Hydrogenation at the N-centre (step 19) is mildly endergonic by 0.05 eV with an activation barrier of 0.89 eV, whereas hydrogenation at the O-centre (step 20) is mildly endergonic by 0.08 eV with an activation barrier of 0.24 eV. Interestingly, hydrogenation becomes more and more favourable as the number of hydrogens increases.

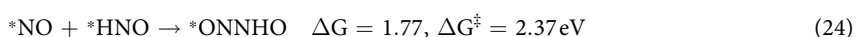
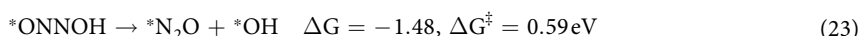
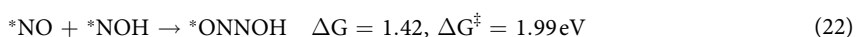


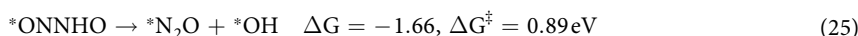
The N-O dissociation from $*H_2NOH$ is highly exergonic (−2.14 eV, step 21) with an activation barrier of 0.32 eV. However, the N-O bond dissociation barrier is lower in $*HN-OH$ (0.14 eV) than in $*H_2N-O$, indicating that $*HNOH$ is likely to dissociate (step 17). Therefore, the N-O bond dissociation barriers are lower and lower for the higher hydrogenated intermediates.

The N-O bond distances in the intermediate species support such a trend. The calculated N-O bond distances are 1.22 Å, 1.40 Å, 1.43 Å and 1.44 Å in $*NO$, $*NOH$, $*HNOH$ and $*H_2NOH$ respectively, suggesting that N-O bond dissociation is very likely for the higher hydrogenated products.

After going through all the possible hydrogenation and non-hydrogenation pathways, we find N_2 formation is highly favourable over NH_3 formation. Therefore, the most favourable route toward N_2 formation is $*NO \rightarrow *N \rightarrow *N_2$, whereas $*NO \rightarrow *ONNH \rightarrow *N_2O \rightarrow *N_2$ is the less favourable route. The most favourable route toward NH_3 formation is $*NO \rightarrow *NOH \rightarrow *HNOH \rightarrow *NH \rightarrow *NH_2 \rightarrow *NH_3$ and the less favourable routes are $*NO \rightarrow *HNO \rightarrow *HNOH \rightarrow *NH \rightarrow *NH_2 \rightarrow *NH_3$ and $*NO \rightarrow *HNO \rightarrow *HNOH \rightarrow *H_2NOH \rightarrow *NH_2 \rightarrow *NH_3$.

Non-favourable Pathways. Along with the most favourable pathways, we have studied some other possible pathways, which are discussed as follows.





$^*\text{NOH}$ could react with $^*\text{NO}$ to form $^*\text{ONNOH}$ (step 22) with an activation barrier of 1.99 eV. As the barrier is very high, this reaction is very unlikely to proceed. $^*\text{ONNOH}$ has a dissociation barrier of 0.59 eV for the formation of $^*\text{N}_2\text{O}$ (step 23). The other possibility is $^*\text{NO}$ reacting with $^*\text{HNO}$ for the formation of a N-N bond (step 24; via $^*\text{ONNH}\text{O}$) followed by N-O bond dissociation (step 25). The N-N bond formation step is highly endergonic by 1.77 eV with an activation barrier of 2.37 eV, and thus is very unlikely to happen. However, the N-O bond dissociation (step 25) is highly exergonic (-1.66 eV) with an activation barrier of 0.89 eV. The calculated activation barriers are 2.37 eV and 0.89 eV for N-N bond formation and N-O bond dissociation steps respectively. The lower activation barrier for the formation of $^*\text{ONNOH}$ over $^*\text{ONNH}\text{O}$ is due to the stability (by 0.15 eV) and strong adsorption energy (by 2.58 eV) of $^*\text{ONNOH}$ species compared with $^*\text{ONNH}\text{O}$. The lower activation barrier for N-O bond dissociation in $^*\text{ONNOH}$ over $^*\text{ONNH}\text{O}$ is attributed to the longer N-O bond distance (1.42 Å) in $^*\text{ONNOH}$ than in $^*\text{ONNH}\text{O}$ (1.32 Å). Such higher activation barrier for $^*\text{ONNH}\text{O}$ (2.04 eV) formation over $^*\text{ONNOH}$ formation (1.42 eV) has been reported previously on the Pt(111) surface⁷.

Kinetic Analysis. From the elementary pathways, we discover that many pathways are possible for a N-O bond dissociation reaction. The Gibbs free energy versus reaction coordinate gives an overall idea to locate the minimum energy pathway from several possibilities. The roles of surface coverage, gas phase partial pressures of reactant and product and reaction temperature cannot be fully understood from the Gibbs free energy calculations. These experimental parameters can provide further insights to find out the most favoured pathway for the NO reduction reaction. Thus, we have done a detailed microkinetic analysis based on our preliminary DFT results to understand the roles of surface coverage, partial pressures of reactants/products and reaction temperatures toward the reaction kinetics. The forward (k_i) and backward (k_{-i}) rate constants for all the elementary steps are calculated using the following equation:

$$k_{i/-i} = \left(\frac{k_B T}{h} \right) \left(\frac{q_F}{q_I} \right) e^{-\Delta G^\ddagger/k_B T} \quad (26)$$

where k_B is the Boltzmann constant, T is the temperature, h is the Planck constant. Here q_i and q_f are vibrational partition functions for the initial and final state structures and ΔG^\ddagger is the Gibbs free energy barrier for the initial and final state of the elementary reaction. The vibrational partition functions (q) are calculated using the following equation:

$$q = \sum_i \frac{1}{1 - e^{-h\nu_i/k_B T}} \quad (27)$$

where ν_i are the vibrational frequencies. All the exergonic reactions are assumed to be irreversible. Hence only forward steps are considered. For endergonic reactions, both the forward and backward steps are considered while developing the microkinetic model. k_i and k_{-i} are used for the forward and backward reaction steps respectively for the i -th step. The details of the microkinetic model are given in the Supporting Information.

Here, using the microkinetic model, we would like to understand the role of hydrogen partial pressure toward the hydrogenation of NO. In the presence of high hydrogen partial pressures, we would like to study whether or not $^*\text{NO}$ will favour hydrogenation over $^*\text{N-O}$ direct dissociation.

As the experimental NO reduction temperature ranges from 300 K to 500 K^{68,69}, the rate constants (Table 1) are calculated in the same temperature range (300 K to 500 K). The rate constants improve significantly as we increase the temperature. At 300 K, the ratio between the rate constants (k_{10}/k_{11}) is $\sim 10^9$ for $^*\text{NO}$ hydrogenation at O- and N-sites respectively. Hence, $^*\text{NOH}$ formation is highly favourable over $^*\text{HNO}$ formation ($^*\text{NO} + ^*\text{H} \rightarrow ^*\text{NOH}$ and $^*\text{NO} + ^*\text{H} \rightarrow ^*\text{HNO}$). The higher possibility of $^*\text{NOH}$ formation can be further confirmed from the surface coverage (Θ) study. The surface coverage ratio of the $\Theta_{\text{NOH}}/\Theta_{\text{HNO}}$ (see Supporting Information for details) is calculated using the steady state approximation⁷⁰, where the rates of production and decomposition are assumed to be equal. Therefore, Θ_{NOH} and Θ_{HNO} are calculated considering their equal rate of formation and decomposition in the complete reaction pathway. The $\Theta_{\text{NOH}}/\Theta_{\text{HNO}}$ ratio is calculated to be 1.78, suggesting that the hydrogenation of $^*\text{NO}$ proceeds through the NOH intermediate, rather than through the $^*\text{HNO}$ intermediate. Surprisingly, the ratio between the forward and backward rate constants (k_{10}/k_{11} ; $\sim 10^9$) is very high compared to their coverage ratio (1.78). The large difference is due to the high dissociation kinetics of $^*\text{NOH}$ (k_{-10}) compared with $^*\text{HNO}$ (k_{-11}). This is very much consistent with the previous trend observed from the relative energetics and activation barriers study, where we find hydrogenation at the O-centre is preferred over the N-centre.

However, hydrogenation at the N- and O- (of NO) is very important as hydrogenation at $^*\text{N}$ and O^* sites will favour the formation of $^*\text{NH}_3$ and $^*\text{N}_2$ respectively, which in turn decides the product selectivity. The dissociation of $^*\text{NOH}$ leads to the formation of $^*\text{N}$, whereas $^*\text{HNO}$ leads to the formation of $^*\text{NH}$. $^*\text{NH}$ can be further hydrogenated for the formation of $^*\text{NH}_3$, whereas $^*\text{N}$ can undergo either hydrogenation or nitrogeneration toward the formation of $^*\text{NH}_3$ and $^*\text{N}_2$ respectively. Thus, it is very important to understand that these hydrogenated pathways improve the product ($^*\text{N}_2$ vs. $^*\text{NH}_3$) selectivity.

Here $\Theta_{\text{NOH}}/\Theta_{\text{HNO}}$ ratio implies the partial pressure ratio of NO and H_2 . Several ratios of partial pressures of NO/H_2 are used to understand their roles toward the product selectivity. The surface coverage ratios of $\Theta_{\text{NOH}}/\Theta_{\text{HNO}}$ are calculated for different partial pressure ratios of $p_{\text{NO}}/p_{\text{H}_2}$. The calculated $\Theta_{\text{NOH}}/\Theta_{\text{HNO}}$ are 5.53×10^{-02} , 1.59×10^{-01} , 3.13×10^{-01} , 6.41×10^{-01} , 1.77×10^{00} , 4.98×10^{00} , 1.22×10^{01} , 3.51×10^{01} and 3.92×10^{02} for 500, 40, 12, 4, 1, 0.25, 0.08, 0.025, 0.002 atmospheres of partial pressure ($p_{\text{NO}}/p_{\text{H}_2}$) respectively.

Elementary reactions	300 K		350 K	400 K	450 K	500 K
	k_f	k_{-f}	k_f	k_f	k_f	k_f
*NO → *N + *O (k_1)	4.77×10^{-05}	3.88×10^{-17}	1.57×10^{-02}	1.23×10^{00}	3.72×10^{01}	5.76×10^{02}
*N + *N → *N ₂ (k_2)	5.07×10^{-06}	1.24×10^{-08}	1.89×10^{-03}	1.63×10^{-01}	5.28×10^{00}	8.65×10^{01}
*NO + *N → *N ₂ O (k_3)	1.39×10^{-30}	3.68×10^{-03}	1.98×10^{-24}	8.27×10^{-20}	3.30×10^{-16}	2.54×10^{-13}
*N ₂ O → *N ₂ + *O (k_4)	2.39×10^{06}	9.11×10^{-38}	2.64×10^{07}	1.63×10^{08}	6.80×10^{08}	2.16×10^{09}
*N + *H → *NH (k_5)	8.02×10^{03}	3.02×10^{02}	1.49×10^{05}	1.36×10^{06}	7.69×10^{06}	3.11×10^{07}
*NH + *H → *NH ₂ (k_6)	7.42×10^{-05}	2.46×10^{04}	2.32×10^{-02}	1.76×10^{00}	5.16×10^{01}	7.79×10^{02}
*NH ₂ + *H → *NH ₃ (k_7)	3.68×10^{04}	3.53×10^{04}	6.33×10^{05}	5.44×10^{06}	2.94×10^{07}	1.15×10^{08}
*NO + *NH → *ONNH (k_8)	4.41×10^{-11}	1.52×10^{13}	1.14×10^{-07}	4.20×10^{-05}	4.22×10^{-03}	1.71×10^0
*ONNH → *N ₂ O + *H (k_9)	2.70×10^{-08}	1.95×10^{-07}	2.78×10^{-05}	5.87×10^{-03}	3.02×10^{-01}	7.94×10^{00}
*NO + *H → *NOH (k_{10})	1.99×10^{-09}	2.08×10^{07}	2.83×10^{-06}	6.66×10^{-04}	4.72×10^{-02}	1.44×10^{00}
*NO + *H → *HNO (k_{11})	1.51×10^{-18}	2.35×10^{00}	4.36×10^{-14}	9.82×10^{-11}	4.04×10^{-08}	5.03×10^{-06}
*NOH → *N + *OH (k_{12})	1.25×10^{-02}	2.69×10^{-24}	1.97×10^{00}	8.91×10^{01}	1.75×10^{03}	1.92×10^{04}
*HNO → *NH + *O (k_{13})	1.08×10^{00}	4.31×10^{-31}	7.59×10^{01}	1.87×10^{03}	2.30×10^{04}	1.73×10^{05}
*NOH + *H → *HNOH (k_{14})	9.18×10^{-09}	1.16×10^{04}	9.90×10^{-06}	1.90×10^{-03}	1.15×10^{-01}	3.08×10^{00}
*HNO + *H → *HNOH (k_{15})	6.59×10^{03}	6.28×10^{08}	1.41×10^{05}	1.43×10^{06}	8.75×10^{06}	3.78×10^{07}
*HNO + *H → *H ₂ NO (k_{16})	2.68×10^{-03}	8.19×10^{02}	4.73×10^{-01}	2.33×10^{01}	4.91×10^{02}	5.67×10^{03}
*HNOH → *NH + *OH (k_{17})	3.82×10^{10}	1.62×10^{-20}	1.00×10^{11}	2.11×10^{11}	3.81×10^{11}	6.18×10^{11}
*H ₂ NO → *NH ₂ + *O (k_{18})	1.79×10^{09}	8.28×10^{-21}	6.86×10^{09}	1.91×10^{10}	4.31×10^{10}	8.34×10^{10}
*HNOH + *H → *H ₂ NOH (k_{19})	7.11×10^{-03}	6.57×10^{-02}	1.12×10^{00}	5.05×10^{01}	9.91×10^{02}	1.09×10^{04}
*H ₂ NO + *H → *H ₂ NOH (k_{20})	4.66×10^{08}	1.22×10^{10}	2.08×10^{09}	6.49×10^{09}	1.59×10^{10}	3.31×10^{10}
*H ₂ NOH → *NH ₂ + *OH (k_{21})	3.29×10^{07}	3.53×10^{-29}	2.19×10^{08}	9.27×10^{08}	2.89×10^{09}	7.24×10^{09}
*NO + *NOH → *ONNOH (k_{22})	2.18×10^{-21}	1.54×10^0	1.55×10^{-16}	6.90×10^{-13}	8.06×10^{-03}	9.13×10^{-08}
*ONNOH → *N ₂ O + *OH (k_{23})	8.59×10^{02}	1.13×10^{-22}	2.61×10^{04}	3.43×10^{05}	2.58×10^{06}	1.31×10^{07}
*NO + *HNO → *ONNHO (k_{24})	8.48×10^{-28}	5.70×10^{02}	4.86×10^{-22}	1.03×10^{-17}	2.42×10^{-14}	1.22×10^{-11}
*ONNHO → *N ₂ O + *OH (k_{25})	8.06×10^{-03}	1.00×10^{-30}	1.28×10^{00}	5.85×10^{01}	1.16×10^{03}	1.28×10^{04}

Table 1. Rate constants (s^{-1}) of the elementary reactions at different temperatures and here k_f and k_{-f} are for the forward and backward steps respectively.

Therefore, *NOH coverage increases with increasing partial pressure of hydrogen, which further improves the product selectivity (*N₂ formation).

From the DFT calculated parameters (Gibbs free energy and activation barriers) and the rate constant of the reaction analysis, we have already proposed that *N₂ is the predominant product over *NH₃ in the presence and absence of hydrogen. We find the spontaneous formation of *N₂ with lower activation barrier and the exergonic behaviour of *NH₂ dissociation with low activation barrier are the underlying reasons for the *N₂ formation with higher selectivity. However, we wonder whether increasing hydrogen partial pressures can turn the selectivity toward *NH₃? Therefore, after cropping the crucial steps from the complete microkinetic model, we have shown the effect of external hydrogen partial pressure toward the product selectivity. The small and effective microkinetic model is developed considering the following steps: (i) *NO + * $\xrightarrow{k_1}$ *N + *O, (ii) *N + *N $\xrightarrow{k_2}$ N₂(g) + 2*, (iii) *N + *H $\xrightarrow{k_5}$ *NH + *, and (iv) *NH + *H $\xrightarrow{k_6}$ *NH₂ + *.

Therefore, the rate of formation of N₂ and NH₂ ($\frac{d\Theta_{N_2}/dt}{d\Theta_{NH_2}/dt}$) can be written as follows (see Supporting Information for details):

$$\frac{d\Theta_{N_2}/dt}{d\Theta_{NH_2}/dt} = \left(\frac{k_1}{k_5}\right) \left(\frac{K_{NO}}{K_{H_2}}\right) \left(\frac{p_{NO}}{p_{H_2}^{\frac{1}{2}}}\right) - 1 \quad (28)$$

Several ratios of partial pressures of NO/H₂ are used to understand their role toward the product selectivity. The calculated ($\frac{d\Theta_{N_2}/dt}{d\Theta_{NH_2}/dt}$) value is $\approx 6.77 \times 10^{35}$ when the partial pressure ratio (p_{H_2}/p_{NO}) of 500:1 is used. Therefore, our results suggest that even with extreme hydrogen partial pressures, N₂ is the predominant product.

However, under the 1:1 partial pressure ratio of NO and H₂, $\Theta_{NOH}/\Theta_{HNO}$ is found to be 1.78. Therefore, there could be a possibility of the reaction proceeding through the *HNO intermediate. Therefore, it is necessary to examine the extent of possibility for proceeding the reaction further from the *HNO intermediate. The ratio of rate constants (k_2/k_5) for (*N $\xrightarrow{k_5}$ *N + *H) and *N + *N $\xrightarrow{k_2}$ *N₂ is 1.58×10^{09} , implying that the *NH formation is favourable over *N₂. However, the ratio of rate constants (k_{-5}/k_6) for *NH dissociation (*NH $\xrightarrow{k_{-5}}$ *N + *H) and *NH₂ formation (*NH + *H $\xrightarrow{k_6}$ *NH₂) is 4.17×10^{06} , suggesting the tendency of *NH to decompose to *N instead of hydrogenation to *NH₂. Therefore, it does not matter whether the reaction proceeds through the *HNO

intermediate or not as it dissociates into *NH and *O (*HNO \rightarrow *NH + *O). Interestingly, further hydrogenation on *NH is highly unfavourable (*NH + *H \rightarrow *NH₂) compared to *NH decomposition (*NH \rightarrow *N + *H). Therefore, the Ni₈₅ nanocluster shows excellent catalytic activity toward high product (N₂) selectivity.

In the earlier sections, we have shown that *NOH coverage increases with increasing partial pressure of hydrogen. This in fact improves the formation of *N₂. But with increasing of hydrogen partial pressures, does the formation of higher hydrogenated products (*HNOH and *H₂NO) increase or not? Moreover, do such higher hydrogenated products also show greater selectivity towards the *N₂ formation over *NH₃ formation or not? The rate constants ratio [(k₁₄ + k₁₅)/k₁₆] for the formation of *HNOH (*NOH + *H \rightarrow *HNOH, *HNO + *H \rightarrow *HNOH) and *H₂NO is 5.88 \times 10⁰⁶, implying that *HNOH is the major product. Now, after N-O bond dissociation, *HNOH and *H₂NO lead to the formation of *HN and *NH₂, respectively. Interestingly, we already found *NH favours the formation of *N₂ over *NH₃.

Therefore, our study shows that the Ni₈₅ nanocluster can efficiently and selectively reduce nitric oxide toward nitrogen even in the presence of high hydrogen partial pressure. Moreover, the most favourable pathway for *N₂ formation is *NO \rightarrow *N \rightarrow *N₂, whereas the least favourable pathway for *N₂ formation is *NO \rightarrow *ONNH \rightarrow *N₂O \rightarrow *N₂. Our results on the Ni-nanocluster are very much promising for product selectivity compared to previous experimental and theoretical reports on noble metal based catalysts (Pt, Pd, Rh, Ag, Pd).

Discussions

DFT calculations have been performed to understand the nitric oxide reduction reaction over (111) facet of the octahedral nickel nanocluster (Ni₈₅) enclosed by well-defined facets. Our results on the Ni-nanocluster show that direct N-O bond dissociation is thermodynamically as well as kinetically very much favourable even in the presence of hydrogen, which is completely opposite to earlier experimental and theoretical reports on noble metal based catalysts (Pt, Pd, Rh, Ag), where indirect N-O bond dissociation (from their hydrogenated products: NOH, HNO, HNOH, H₂NO) is favourable over direct N-O bond dissociation. Hence, the product selectivity (N₂ vs. N₂O/NH₃) increases on a Ni-nanocluster surface, whereas selectivity decreases on noble metal surfaces. The NO reduction reaction on Ni(111) bulk surfaces is reported by several groups, but no one has studied the reduction reaction in the presence of the hydrogen. Thus, a vis-à-vis comparison could not be made. The lower *N-O bond dissociation and N-N bond formation barriers make the Ni₈₅ cluster an efficient and a selective catalyst toward NO reduction for N₂ formation. In the presence of hydrogen, the reaction might proceed through *NOH or higher hydrogenated intermediates, but such intermediates easily dissociate into *NH or *N, which in turn favours *N₂ formation over *NH₃. Our microkinetic analysis shows that even with extreme hydrogen partial pressures, *N₂ formation is favourable over *NH₃. Thus, the product selectivity towards N₂ does not change even under high hydrogen partial pressures on a Ni-nanocluster surface, which is again completely opposite to earlier experimental and theoretical reports on noble metal based catalysts (Pt, Pd, Rh, Ag). Hence, we report that nickel based nanoclusters could be very promising catalysts for the efficient and selective reduction of nitric oxide to nitrogen.

Methods

The first-principles calculations are performed using the projected augmented wave (PAW) method⁷¹, as implemented in the Vienna ab initio simulation package (VASP)⁷²⁻⁷⁴. The exchange-correlation potential is described by using the generalized gradient approximation of Perdew-Burke-Ernzerhof (GGA-PBE)⁷⁵. The projector augmented wave (PAW) method⁷¹ is employed to treat interactions between ion cores and valence electrons. The structure optimization is based on the conjugate gradient-minimization scheme under a spin polarization consideration. Our calculated magnetic value shows that each nickel has a magnetic moment of 0.71 μ_B which is in very much agreement with the previously reported value of 0.69 μ_B ²⁵. For the consideration of van der Waals interactions, we have used Grimme's D3-type of semiempirical method for dispersion energy correction^{76,77}. But, we have listed the non-dispersion corrected values for a vis-à-vis comparison with earlier reports, as they have not included dispersion correction in their calculations. However, we have given both the dispersion-corrected and non-corrected values in the Supporting Information. A 22 \times 22 \times 22 Å³ cubic supercell is used to optimize the metal clusters to rule out the possibility of interaction of periodically repeated clusters. The Brillouin zone is sampled using gamma k-point (1 \times 1 \times 1). The total energy improved by 0.0001 eV when k-points are increased to 2 \times 2 \times 2. As the box size is quite high therefore we have used gamma point for our calculations. Plane wave with a kinetic energy cut off of 470 eV is used to expand the electronic wave functions. All the atoms are relaxed for the full structural relaxation. The climbing nudged elastic band (CI-NEB) method⁷⁸ is used to locate the transition states. Six intermediate images are used in each CI-NEB pathway. Vibrational frequencies for the initial, transition and final states of the reactions are calculated and the transition state is confirmed from the presence of one imaginary frequency along the reaction coordinate. Zero-point energy (ZPE) is calculated from the following equation:

$$ZPE = \sum_i 1/2 h v_i$$

where h is the Planck constant and v_i is the frequency of the i^{th} vibrational mode.

The reaction free energies (ΔG) and activation barriers (ΔG^\ddagger) are calculated using zero point energy (ZPE) and entropy corrections. The species are denoted with asterisk (*) while adsorbed on the surface.

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Author Contributions

A.M. designed the project, did the calculation work and wrote the manuscript. K.S.R. and I.C. helped in the scientific discussion and commented on the manuscript. B.P. supervised the project and co-wrote the manuscript.

Additional Information

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